**1/30/15 Basic Chemistry**

**Blood**

**1. What compounds are in blood?**

**2. Why is blood a mixture?**

**3. What is the acceptable pH range for blood?**

**4. How is this maintained?**

**5. When pH levels lower dangerously, acidosis results. This**

**happens when acids build up or bicarbonate ions are lost.**

**Give causes for each of the following reasons acids build up:**

1. **too much carbon dioxide (ie. Can’t get rid of enough CO2)**
2. **kidney can’t remove acids (ketoacidosis)**
3. **lactic acid build up**

**Give a cause for losing too much sodium bicarbonate**

**(hyperchloremic acidosis)**

**6. How does the body react when acidosis results?**

**7. When pH levels rise dangerously, alkalosis results. Give causes**

**for the following reasons of alkalosis:**

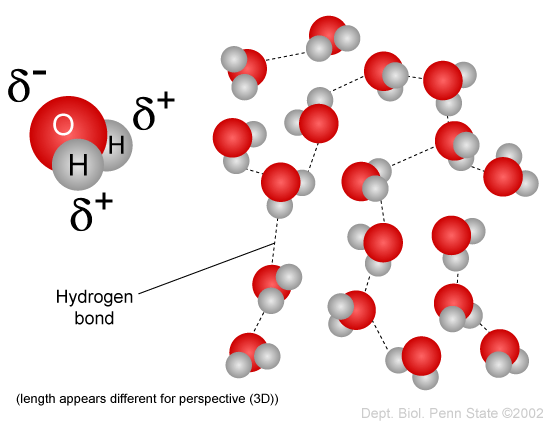
1. **too little CO2**
2. **too little chlorine**

**\*c. lack of oxygen = high altitudes**

**8. How does the body react when alkalosis results?**

**Biochemistry: Inorganic and Organic Chemistry**

1. **Inorganic**
2. **Water –**

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1. **Salts –**
2. **Acids – already in notes**
3. **Bases – already in notes**
4. **Oxygen – already in notes**